# Hydroxyl-Directed Cross-Coupling: A Scalable Synthesis of Debromohamigeran E and Other Targets of Interest

Thomas P. Blaisdell and James P. Morken*

Department of Chemistry, Merkert Chemistry Center, Boston College, Chestnut Hill, Massachusetts 02467, United States

**S** *Supporting Information*

**ABSTRACT:** A hydroxyl functional group positioned $\beta$ to a pinacol boronate can serve to direct palladium-catalyzed cross-coupling reactions. This feature can be used to control the reaction site in multiply borylated substrates and can activate boronates for reaction that would otherwise be unreactive.

The stereoselective diboration reaction is a versatile approach for the efficient transformation of olefin substrates: the vicinal 1,2-bis(boronate) products possess reactivity that facilitates functionalization of both carbons of the original alkene framework.[1] While oxidation of vicinal bis(boronates) to diols is broadly applicable to an array of substrates, differential transformation of the two boronates requires a substrate bias that allows site-selective transformation of one boronate in preference to the other. In this context, the sensitivity of the Suzuki−Miyaura cross-coupling to steric effects allows selective transformation of a primary, terminal boronate without detectable reaction of an adjacent secondary boronate group. Our group has employed this feature in a selective tandem diboration/cross-coupling (DCC) strategy that allows the transformation of terminal alkenes into an array of functionalized nonracemic products via **A** (Scheme 1).[2] It was considered that a

**Scheme 1. Site Selectivity in Cross-Coupling of Vicinal Bis(boronates)**
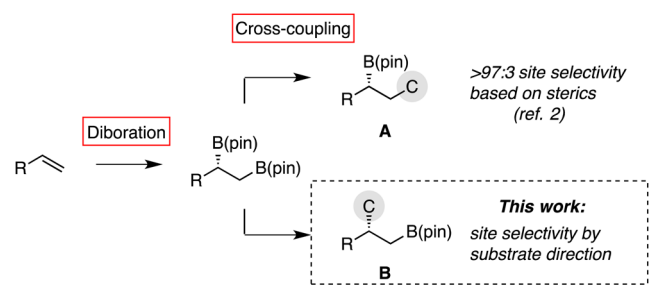


different array of target structures would be accessible if the Suzuki−Miyaura cross-coupling could be coaxed to occur preferentially at the secondary boronate in the substrate to give **B**.[3] Importantly, features that overturn the steric bias present in terminal bis(boronates) might also enable the site-selective cross-coupling of bis(boronates) derived from internal alkene substrates. Here, we show that a hydroxyl group can control the site-selectivity of a Suzuki−Miyaura reaction. A noteworthy feature of this process is that it appears to rely upon a substrate-

directed transmetalation that may involve a covalent linkage between the directing hydroxyl group and the Pd center. This reactivity mode is not commonly employed to facilitate catalytic cross-coupling reactions[4] and may be a useful element in reaction design.
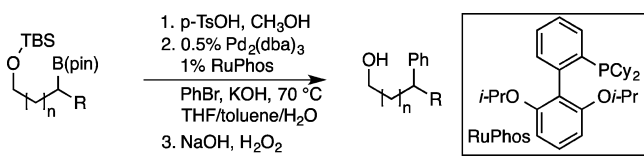
One strategy for the stereospecific cross-coupling of secondary alkyl boronates employs Lewis basic functional groups to enhance the reactivity of an adjacent boronic acid derivative. In this context, Suginome,[5] Molander,[6] Hall,[7] and Takacs[8] have employed adjacent amide or ester functional groups as electron donors to facilitate cross-coupling. Considering that protected alcohols are readily accommodated in asymmetric diboration, and since alcohols are an inherent feature in directed metal-free diboration reactions,[9] we probed the ability of a hydroxyl group to control regioselectivity of cross-coupling (Table 1). In these experiments, purified borylated alkyl silyl ethers were first subjected to TBS deprotection, then reagents for a Pd/RuPhos[10] mediated cross-coupling were introduced. After reaction for 12 h at 70 °C, the reaction mixture was subjected to oxidative workup. Preliminary experiments (see Supporting Information) indicated that the reaction solvent plays a critical role in controlling the ratio of elimination products to cross-coupling products with a THF/toluene/$H_2O$ combination providing optimal selectivity.
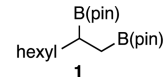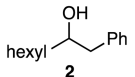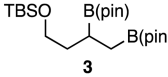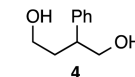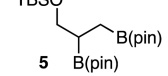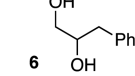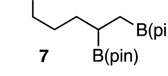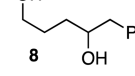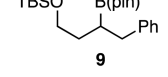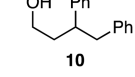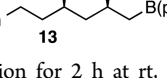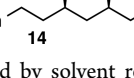
When the above-described cross-coupling conditions were applied to nonfunctionalized bis(boronate) **1**, secondary alcohol **2** was furnished in 71% yield. In contrast, when bis(boronate) **3** was subjected to the deprotection/cross-coupling/oxidation sequence, the presence of the hydroxyl group results in a complete turnover in regioselectivity such that **4** was provided in 72% yield. Structures **9**, **11**, and **12** provide clear evidence that while directing effects from the $\beta$-position are efficient (entry 5), when the directing group is positioned at the $\alpha$ or $\gamma$ site (entries 6 and 7) the cross-coupling does not benefit from the presence of a neighboring hydroxyl group.

While the ability of a $\beta$ hydroxyl group to activate a secondary pinacol boronate for cross-coupling is reminiscent of the activation provided by $\beta$-acyl groups, the reactions in Table 1 appear to operate by a distinct mechanistic principle relative to the 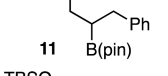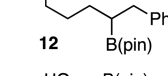acyl-promoted cross-couplings developed by Suginome, Molander, and Hall.[5−7] This distinction is most clearly indicated by the observation that acyl-promoted cross-couplings proceed with inversion of configuration, ostensibly by generalized chelated "ate" complex **C** (Scheme 2a), whereas the hydroxyl-directed cross-coupling occurs with complete retention of configuration at carbon (Table 1, entry 8). The stereochemical

**Table 1. Survey of Substrates for Pd/RuPhos Catalyzed Cross-Coupling of Alkylboronates and Bromobenzene**[a]



| entry | substrate | product | % yield[b] |
|---|---|---|---|
| 1 | **1** | **2** | 71 |
| 2 | **3** | **4** | 72 |
| 3 | **5** | **6** | 77 |
| 4 | **7** | **8** | 29 |
| 5 | **9** | **10** | 70 |
| 6 | **11** | no reaction | - |
| 7 | **12** | no reaction | - |
| 8[c] | **13** | **14** | 67 |

[a]Deprotection for 2 h at rt, followed by solvent removal and cross-coupling for 12 h at 70 °C. [b]Yield determined by ^{1}H NMR versus 1,3,5-trimethoxybenzene as an internal standard. [c]The substrate was prepared by $Cs_2CO_3$-catalyzed diboration in THF/methanol and the solvent evaporated prior to subjection to cross-coupling.

outcome with hydroxyl direction is reminiscent of Takacs' secondary-amide-directed stereoretentive cross-coupling, although the mechanistic underpinnings for this process are unresolved.[8] Additional mechanistic information about the hydroxyl-directed reaction was obtained by examining the cross-coupling of bis(boronate) **15** (Scheme 2b). In line with the observations in Table 1, when **15** was subjected to Pd/RuPhos catalyzed coupling with 1.05 equiv of bromobenzene, the reaction exhibited complete selectivity for the B(pin) group positioned $\beta$ to the hydroxyl substituent. Of note, in addition to cross-coupling product **16**, ketones **17** and **18**, and unreacted **15** were isolated. When the reaction was conducted in the presence of excess bromobenzene, complete conversion of **15** occurs, and only **16** and **17** were isolated as the reaction products. Considering the known capacity for Pd(II) complexes to oxidize alcohols with aryl halides as terminal oxidant,[11] a plausible rationale for formation of **17** and **18** involves $\beta$-hydrogen elimination from a substrate-derived L(Ar)Pd(alkoxide). Thus, the observation of products **17** and **18** establishes the capacity for substrate-derived Pd(alkoxide) formation during the course of catalytic cross-coupling reactions. A last piece of data pertaining to the origin of hydroxyl-direction arises from the X-ray structure

**Scheme 2. Features of Directed Cross-Coupling Reactions**
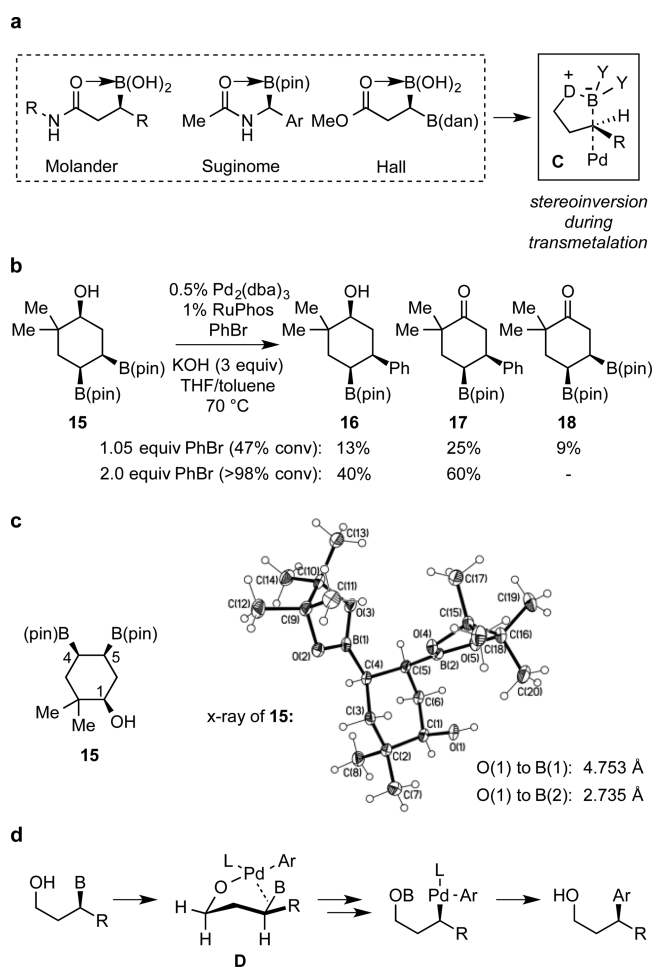


of **15** (Scheme 2c). In the solid-state structure, both the $\beta$ B(pin) and the hydroxyl group occupy axial positions; however, the planar geometry of boron suggests that the two groups are too far removed for internal HO → B coordination. Thus, even in a case where "ate" complex formation appears to be precluded, the cross-coupling is site selective. Taking the above observations into account, a plausible mechanism for the directing effect observed in Table 1 involves binding of the substrate hydroxyl to an LPdAr complex, perhaps by displacement of a halide; subsequent internal delivery of Pd through a complex such as **D** (Scheme 2d) would generate an organopalladium complex through an inner-sphere stereoretentive transmetalation and ultimately deliver the corresponding coupling product.

The hydroxyl-directed transmetalation can merge seamlessly with recently developed hydroxyl-directed metal-free diboration and allows for rapid stereoselective and site-selective functionalization of homoallylic alcohols. As depicted in Scheme 3, sequential directed diboration/cross-coupling, when followed by silylation or acylation, furnishes $\gamma$-oxygenated boronates from a range of substrates. Importantly, the reaction can be conducted with aryl, heteroaryl, and alkenyl electrophiles and applies equally well to terminal alkenes, internal olefins, and trisubstituted alkenes. While the terminal and internal olefins furnish 1,3-syn relative stereochemistry, the trisubstituted alkene furnishes modest 1,3-anti induction, an outcome that is inline with the stereoinduction observed in the directed diboration reactions. Lastly, the use of Pd(OAc)$_2$ in reactions of alkenyl electrophiles appeared to minimize alkene-containing side-products derived
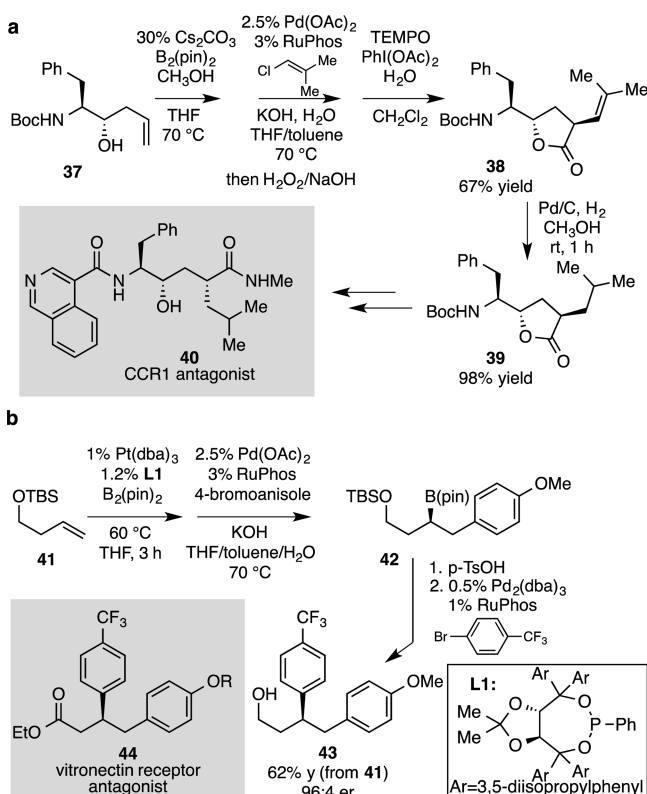
### Scheme 3. Tandem Directed Diboration/Directed Cross-Coupling[a]
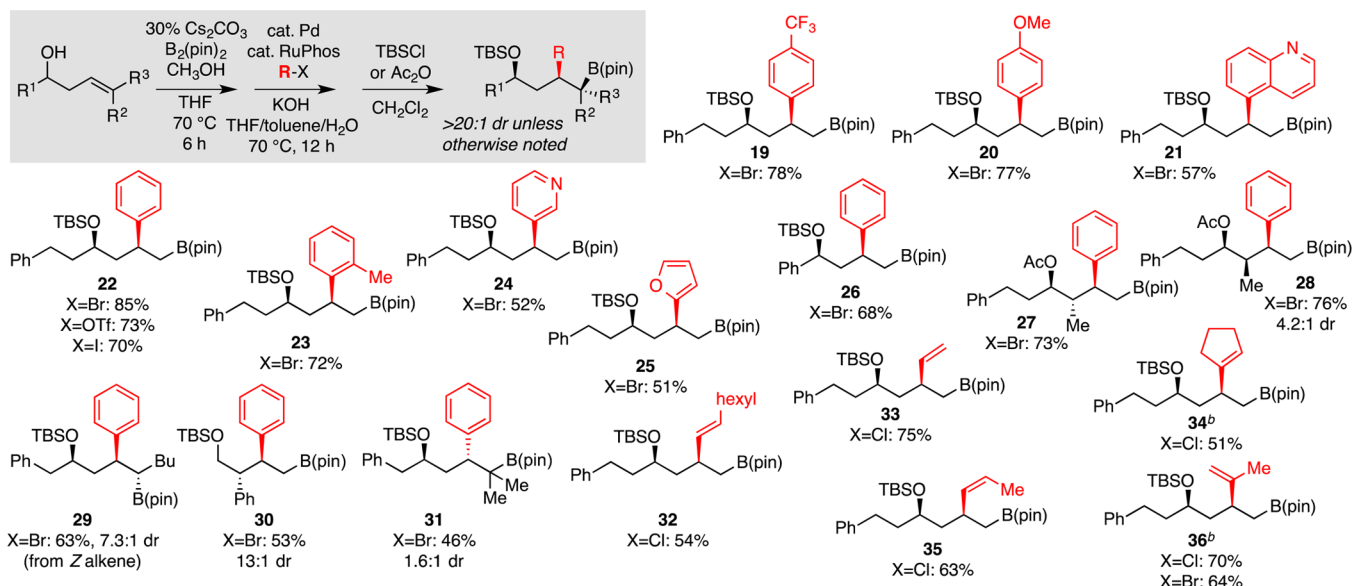


[a]Cross-coupling: aryl electrophiles employed 0.5% Pd₂(dba)₃, 1% RuPhos, and 1.5 equiv ArX; alkenyl electrophiles employed 2.5% Pd(OAc)₂, 3% RuPhos and 1.5 equiv alkenyl halide. Yield represents isolated yield of purified material and is an average of two experiments. [b]Cross-coupling employed 10% Pd(OAc)₂ and 12% RuPhos.

from $\beta$-deborylation of intermediate organopalladium complexes.

Tandem diboration/directed cross-coupling reactions can allow alkene functionalization in ways that are not straightforward with current methods. These strategies are illustrated by the examples in Scheme 4. To target CCR1 antagonist **40**,[12] we considered the reaction of homoallylic alcohol **37**, a starting material that is readily assembled by allylation of Boc-protected phenylalanine.[13] Tandem directed diboration/directed cross-coupling allows stereoselective C−C bond formation at the internal alkene carbon of **37**. Upon TEMPO-catalyzed oxidation, this reaction sequence delivers lactone **38**, which is a precursor to key building block **39**.[12] In a related fashion, enantioselective double arylation of alkene **41** (Scheme 4b) can be accomplished with Pt-catalyzed enantioselective diboration to initiate the sequence.[1c] With the alcohol protected as a silyl ether, steric effects dominate the cross-coupling of the intermediate bis-(boronate) such that bond formation occurs at the terminal carbon. Subsequent TBS deprotection and directed cross-coupling furnishes nonracemic alcohol **43**, a plausible intermediate in the synthesis of vitronectin receptor antagonist **44**.[14]

As a last synthesis example, we considered construction of debromohamigeran E (Scheme 5).[15] This structure belongs to a large group of related cyclopentane-containing compounds[16] but has not yet been addressed by laboratory synthesis.[17] We considered that the target might be accessed by directed cis double alkylation of alkene **46**. To effect this transformation, **46**[18] was subjected to diboration followed by directed cross-coupling with **48**. Remarkably, even though the reacting boronate in **47** is at a secondary carbon and resides between quaternary and tertiary carbons, the cross-coupling proceeds stereoselectively and furnishes a single isomer of product. Of note, this reaction was readily accomplished on a preparative scale and was used to furnish 2.7 g of silyl protected product **49**, an intermediate that was characterized by X-ray crystallography. A subsequent Evans−Zweifel olefination[19] with 2-lithiopropene furnished 2.1 g (93% yield) of double alkylation product **50**, a versatile intermediate for the synthesis of a number of hamigerans. After C−C bond installation, a portion of **50** was reduced and subjected to silyl deprotection. Subsequent Ru-catalyzed oxidation[20] occurred at the primary alcohol site and the benzylic methylene thereby furnishing **52**. Saponification afforded the target structure.

### Scheme 4. Utility of Directed Cross-Coupling

**Scheme 5. Synthesis of (±)-Debromohamigeran E**
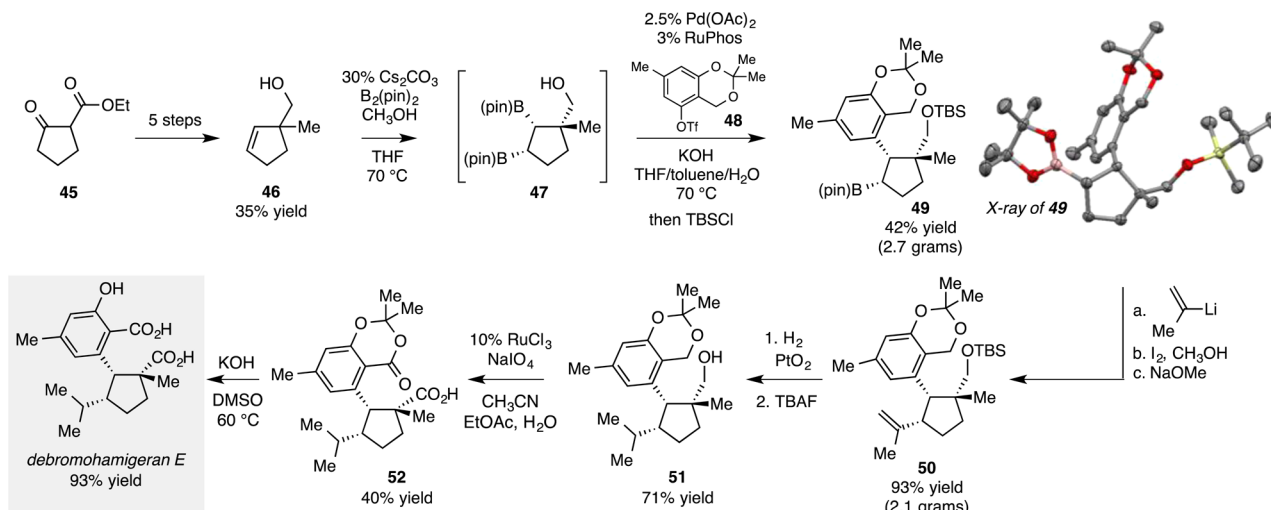


In summary, the hydroxyl-directed cross-coupling enables regioselective, stereoretentive C−C bond formation in the context of multiply borylated reagents. Moreover, the induced proximity conferred by the putative Pd−O linkage appears to facilitate reaction of boronate groups that are likely to be recalcitrant in the absence of such activating effects.

## ASSOCIATED CONTENT

**S** Supporting Information

Procedures, characterization and spectral data. The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/jacs.5b05477.

## AUTHOR INFORMATION

**Corresponding Author**

*morken@bc.edu

**Notes**

The authors declare no competing financial interest.

## ACKNOWLEDGMENTS

## REFERENCES

(1) Reviews: (a) Burks, H. E.; Morken, J. P. *Chem. Commun.* **2007**, 4717. (b) Suginome, M.; Ohmura, T. In *Boronic Acids*, 2nd ed.; **2011**, Wiley-VCH: New York, 2011; Vol. *1*, p 171. (c) Takaya, J.; Iwasawa, N. *ACS Catal.* **2012**, *2*, 1993.

(2) Mlynarski, S. N.; Schuster, C. H.; Morken, J. P. *Nature* **2014**, *505*, 386.

(3) For important recent advances in the cross-coupling of secondary boronates, see: (a) Li, L.; Zhao, S.; Joshi-Pangu, A.; Diane, M.; Biscoe, M. R. *J. Am. Chem. Soc.* **2014**, *136*, 14027. (b) Matthew, S. C.; Glasspoole, B. W.; Eisenberger, P.; Crudden, C. M. *J. Am. Chem. Soc.* **2014**, *136*, 5828. (c) Imao, D.; Glasspoole, B. W.; Laberge, V. S.; Crudden, C. M. *J. Am. Chem. Soc.* **2009**, *131*, 5024. (d) Dreher, S. D.; Dormer, P. G.; Sandrock, D. L.; Molander, G. A. *J. Am. Chem. Soc.* **2008**, *130*, 9257. (e) van der Hoogenband, A.; Lange, J. H. M.; Terpstra, J. W.; Koch, M.; Visser, G. M.; Visser, M.; Korstanje, T. J.; Jastrzebski, J. T. B. H. *Tetrahedron Lett.* **2008**, *49*, 4122. For a review, see: (f) Leonori, D.; Aggarwal, V. K. *Angew. Chem., Int. Ed.* **2015**, *54*, 1082.

(4) For directing effects in Stille couplings, see: (a) Crisp, G. T.; Gebauer, M. G. *Tetrahedron Lett.* **1995**, *36*, 3389. (b) Quayle, P.; Wang, J.; Xu, J.; Urch, C. J. *Tetrahedron Lett.* **1998**, *39*, 485. (c) Domínguez, B.; Pazos, Y.; de Lera, A. R. *J. Org. Chem.* **2000**, *65*, 5917.

(5) (a) Ohmura, T.; Awano, T.; Suginome, M. *J. Am. Chem. Soc.* **2010**, *132*, 13191. (b) Awano, T.; Ohmura, T.; Suginome, M. *J. Am. Chem. Soc.* **2011**, *133*, 20738.

(6) Sandrock, D. L.; Jean-Gérard, L.; Chen, C.; Dreher, S. D.; Molander, G. A. *S J. Am. Chem. Soc.* **2010**, *132*, 17108.

(7) Lee, J. C. H.; McDonald, R.; Hall, D. G. *Nat. Chem.* **2011**, *3*, 894.

(8) Hoang, G. L.; Yang, Z.-D.; Smith, S. M.; Pal, R.; Miska, J. L.; Pérez, D. E.; Pelter, L. S. W.; Zeng, X. C.; Takacs, J. M. *Org. Lett.* **2015**, *17*, 940.

(9) (a) Blaisdell, T. P.; Caya, T. C.; Zhang, L.; Sanz-Marco, A.; Morken, J. P. *J. Am. Chem. Soc.* **2014**, *136*, 9264. For nondirected metal-free diboration, see: (b) Bonet, A.; Pubill-Ulldemolins, C.; Bo, C.; Gulyás, H.; Fernández, E. *Angew. Chem., Int. Ed.* **2011**, *50*, 7158.

(10) Charles, M. D.; Schultz, P.; Buchwald, S. L. *Org. Lett.* **2005**, *7*, 3965.

(11) (a) Tamaru, Y.; Yamamoto, Y.; Yamada, Y.; Yoshida, Z. *Tetrahedron Lett.* **1979**, *20*, 1401. (b) Tamaru, Y.; Yamada, Y.; Inoue, K.; Yamamoto, Y.; Yoshida, Z. *J. Org. Chem.* **1983**, *48*, 1286. (c) Bei, X.; Hagemeyer, A.; Volpe, A.; Saxton, R.; Turner, H.; Guram, A. *J. Org. Chem.* **2004**, *69*, 8626. (d) Muzart, J. *Tetrahedron* **2003**, *59*, 5789.

(12) Kath, J. C.; Brissette, W. H.; Brown, M. F.; Conklyn, M.; DiRico, A. P.; Dorff, P.; Gladue, R. P.; Lillie, B. M.; Lira, P. D.; Mairs, E. N.; Martin, W. H.; McElroy, E. B.; McGlynn, M. A.; Paradis, T. J.; Poss, C. S.; Stock, I. A.; Tylaska, L. A.; Zheng, D. *Bioorg. Med. Chem. Lett.* **2004**, *14*, 2169.

(13) Diaz, L. C.; Diaz, G.; Ferreira, A. A.; Meira, P. R. R.; Ferreira, E. *Synthesis* **2003**, *4*, 603.

(14) Manley, P. J.; Miller, W. H.; Uzinskas, I. N. (SmithKline Beecham Corporation) Patent EP1218005 A2, July 3, 2002.

(15) Isolation: Wellington, K. D.; Cambie, R. C.; Rutledge, P. S.; Bergquist, P. R. *J. Nat. Prod.* **2000**, *63*, 79.

(16) For an overview of the hamigerans, see: Nicolaou, K. C.; Gray, D. L. F.; Tae, J. *J. Am. Chem. Soc.* **2004**, *126*, 613.

(17) For a review of synthesis efforts towards hamigeran B see: Miesch, M.; Welsch, T.; Rietsch, V.; Miesch, L. *Strategies and Tactics in Organic Synthesis*; Harmata, M., Ed.; Academic Press: Oxford, U.K., 2013; Vol 9, p 203

(18) Pichlmair, S.; de Lera Ruiz, M.; Basu, K.; Paquette, L. A. *Tetrahedron* **2006**, *62*, 5178.

(19) (a) Zweifel, G.; Arzoumanian, H.; Whitney, C. C. *J. Am. Chem. Soc.* **1967**, *89*, 3652. (b) Evans, D. A.; Crawford, T. C.; Thomas, R. C.; Walker, J. A. *J. Org. Chem.* **1976**, *41*, 3947.

(20) Bunch, L. (Kobenhavns Universitet) International Patent WO 2014/124651 A1, August 21, 2014.